Level - 1	DTS-5

61.(D) Refer NCERT

62.(B) $SiCl_4 + H_2O \longrightarrow Si(OH)_4 \xrightarrow{\Delta} SiO_2$

63.(B) CO reacts with haemoglobin in blood to form carboxyhaemoglobin.

64.(A) Boiling point of hydrides increases down the group due to increase in Vander Waal radius/molar mass.

65.(D) $\dot{C}H_3$ is sp^2 hybridised. Here unpaired electron is present in pure p-orbital.

66.(C) CO_2 : Bond angle is 180°

SiO_2 exists in form of 3D silicate. Bond angle $= 109.5°$

67.(B) $Na_2B_4O_7 \cdot 10H_2O + HCl \longrightarrow 2NaCl + H_3BO_3$

68.(B) BF_3 is covalent

69.(B) Al_2O_3 is insoluble in water due to high Lattice energy and low hydration energy.

70.(A) $H_2C_2O_4 \xrightarrow[\text{Conc. } H_2SO_4]{\Delta} H_2O + CO + CO_2$

71.(D) $(AlCl_3)_n$ sublimes on heating

72.(B) Number of shared oxygen in sheet silicate is 3.

73.(A) Silicones repel water due to presence of alkyl group

74.(C) $H_2C_2O_4 \longrightarrow \underset{(y)}{CO} + \underset{(x)}{CO_2} + \underset{(z)}{H_2O}$

x : sp , y : sp , z : sp^3

75.(D) Bond order

CO	:	3
CO_2	:	2
CO_3^{2-}	:	1.33

Bond length $\propto \dfrac{1}{\text{Bond order}}$